

NOUN Exams Past Questions

Question

The presence of hydrogen bond in most molecules is responsible for the following except;

Answer

Low boiling point

Question

Which of these is not a covalent hydride?

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NaH

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Hydrogen can be obtained economically as a bi-product in the electrolysis of -----

Answer

brine

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Which of these is not an isotope of hydrogen?

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Which of the isotopes of hydrogen is radioactive?

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Which if these is not a property of hydrogen?

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Hydrogen may not be advantageous as a fuel because-----

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it is a secondary source of energy

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_____types of hydride compounds can be formed depending on the electronegativity of the elements

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The Most significant improvement of mendeleev periodic table came through the discovery of

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Atomic Number

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The law of octaves was rejected for one the following reasons

Answer

mystical connection between music and chemistry

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When two hydrogen atoms combine to form a molecule the two nuclei can spin in the same or opposite direction this phenomena is known as

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Spin isomerism

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What type of hydride is this said to be LiH

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Ionic hydride

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One of the following is a disadvantage of hydrogen

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Why is water a liquid and hydrogen sulphide a gas at room temperature

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Electronegativity of the oxygen atom which polarizes the hydrogen atoms forming stronger hydrogen bonds causing higher boiling points

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Hydrogen bond formed between two atoms of the same or different molecule is called

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Intermolecular hydrogen bonding

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Complete the following reaction $\text{CO} + \text{H}_2\text{O} \rightarrow$

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$\text{CO}_2 + \text{H}_2$

Question

As the principal quantum number increases, the outer electrons get farther away from the nucleus and the atomic radii _____

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Increases

Question

The distance between the nucleus of an ion and the point up to which the nucleus has influence on the electron cloud is called

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One half of the distance between the nuclei of two adjacent metal atoms in the close packed crystal lattice is called

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Half the distance between the nuclei of two atoms belonging to two adjacent molecules in a crystal lattice is called

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Van der waals radius

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One half of the distance between the nuclei of two like atoms bonded together by a single covalent bond is called

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Why do alkali metals not occur as free elements in nature?

Answer

Because they are highly reactive elements

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What principles or rules are violated in the following electronic configuration $(1s^2 2p_x^2)$?

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What is the electronic configuration of elements with atomic number 29?

Answer

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What is the electronic configuration of elements with atomic number 24?

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$\backslash ([\text{Ar}]3d^{\{5\}}4s^{\{1\}}\backslash)$

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The electron affinity decreases down the group why?

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down the group the size of the atom increases and effective nuclear charge remains almost steady

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